

3055 BA SANGAM COLLEGE

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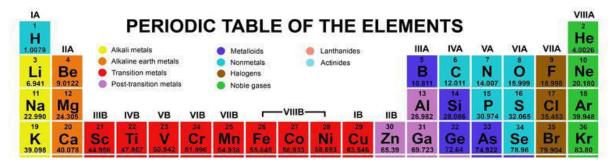
WORKSHEET 8

School: <u>Ba Sangam College</u> Year/level: <u>10</u>

Subject: Basic Science NAME:

Strand	Matter
Sub Strand	Investigating Matter
Content Learning Outcome	Investigate the structure of an atom and explain the properties of common elements in relation to their position on the periodic
	table.

Lesson Notes – *The Periodic Table*



- ❖ **Periodic table** an arrangement of the elements in order of their atomic numbers.
- ❖ Period a horizontal row of elements in the periodic table
 - There are 7 periods e.g. Period 1 elements: H and He

Period 2 elements: Li, Be, B, C, N, O, F and Ne.

- The atomic number increases by one from one element to the next when moving from left to right across a period.
- All elements have the same number of electron shells except the outer shell has one more electron than the previous one.
- ❖ Groups a vertical column of elements in the periodic table.
 - Groups are numbered using Roman numerals one (I) eight (VIII) except transition metals. E.g Group I elements: H, Li, Na and K.(first column of periodic table)
 - Elements in the **same group** have the **same number of electrons** in their **outer shell** and so have **similar chemical properties**.
- Groups with alternative names.
- Group I Alkali Metals Group II Alkali earth Metals
- Group VII Halogens Group VIII Noble gases/Inert

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1. In the periodic table each row is called a	(1 mark)		
2. The elements in each	have the same number of (1 mark)		
3. Each column in the periodic table is called the	(1 mark)		
4. Each of the elements in the same in their outer shells.	have the same number of (2 marks)		
5. Complete the table below.			

Atomic Number	Electron Configuration	Symbol of Element	Group of Periodic Table
3			
12			

(3 marks)

- 6. Write down the name of the elements with the following symbols.
 - i. B _____
 - ii. C
 - iii. N _____
 - iv. O ______ (2 marks)

TOTAL: 10